***Sincerity, Nobility and Service***

**WORKSHEET NO.6**

**GRADE: XII DATE: 19.09.19**

**SUBJECT: CHEMISTRY TOPIC: p-BLOCK ELEMENTS**



1. Why do halogen possess high boiling point?

2. Why do iodine exhibits little metallic character?

3. Why did fluorine exhibits only -1 oxidation states in all its compounds?

4. Why does the tendency of group 16 elements to exists in -2 oxidation state decrease on moving down the group?

5. Why is the first ionization energy of oxygen lower than that of sulphur?

6. Why does oxygen exist in gaseous state where other elements of the group exist in solid state at room temperature?

7. The electron affinity of oxygen is lower than that of sulphur. Explain.

8. Why do halogen atoms have high ionization energies?

9. Why do melting and boiling points of halogens increase in moving down the group?

10. Why does fluorine and oxygen exhibit an anomalous behavior as compared to the other halogens in the group?

11. What type of oxoacids is formed by halogen? Give a brief account of the preparation, important properties and acidic nature of the oxoacids of chlorine.

12. What are interhalogen compounds. Explain.

13. Give the method preparation of:

a. XeF4 b. XeOF4 c. XeF6 d. XeO3

14. Describe the shape of:

a. SiF6 2- b. SiF4 c. PCl5

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